Phys 362
Fall 2019

Statistical and Thermal Physics: Class Exam I

12 March 2019

Name: Solution **Total:** /50

Instructions

- There are 6 questions on 9 pages.
- Show your reasoning and calculations and always explain your answers.

Physical constants and useful formulae

$R = 8.31\,\text{J/mol K}$ $N_A = 6.02 \times 10^{23}\,\text{mol}^{-1}$ $k = 1.38 \times 10^{-23}\,\text{J/K}$ $1\,\text{atm} = 1.01 \times 10^5\,\text{Pa}$

Question 1

A quantity of air, a diatomic gas ($\gamma = \cancel{5/3}$ 7/5), initially occupies a volume of $2.0 \times 10^{-3}\,\text{m}^3$ at atmospheric pressure and at temperature 300 K. This is allowed to expand adiabatically, doubling its volume. Determine the pressure and temperature at the end of this process.

$PV^\gamma = \text{constant} \Rightarrow P_i V_i^\gamma = P_f V_f^\gamma$

$\Rightarrow P_f = P_i \left(\frac{V_i}{V_f}\right)^\gamma = 1\,\text{atm} \left(\frac{1}{2}\right)^{7/5} = 0.38\,\text{atm}.$

Then $PV = NkT$

$\Rightarrow P_f V_f = NkT_f$ and $P_i V_i = NkT_i$

$\Rightarrow \frac{T_f}{T_i} = \frac{P_f V_f}{P_i V_i} = 0.38 \times 2 = 0.76$

$\Rightarrow T_f = 0.76\, T_i$

$= 227\,\text{K}.$

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Question 2

Answer either part a) or part b) for full credit for this problem.

a) Consider a system for which the energy is $E = \alpha V T^4$ and $P = \frac{1}{3}\alpha T^4$. Determine expressions for

$$\left(\frac{\partial E}{\partial V}\right)_P \quad \text{and} \quad \left(\frac{\partial E}{\partial P}\right)_V .$$

Need $E = E(P, V)$.

Then $T^4 = 3P/\alpha$ and so

$$E = \alpha V 3P/\alpha \Rightarrow E = 3VP$$

So $\left(\frac{\partial E}{\partial V}\right)_P = 3P$

$$\left(\frac{\partial E}{\partial P}\right)_V = 3V$$

Question 2 continued ...

b) The enthalpy of a gas is $H = E + PV$. Starting with the infinitesimal version of the first law $\mathrm{d}E = \delta Q + \delta W$ show that the heat capacity at constant pressure is

$$c_P = \left(\frac{\partial H}{\partial T}\right)_P .$$

$$\delta Q = dE - \delta W$$

$$= dE - (-PdV)$$

$$= dE + PdV$$

$$E = H - PV \Rightarrow dE = dH - PdV - VdP$$

So

$$\delta Q = dH - VdP$$

$$= \underbrace{\left(\frac{\partial H}{\partial T}\right)_P}_{C_P} dT + \left[\left(\frac{\partial H}{\partial P}\right)_T - V\right] dP$$

$$\Rightarrow C_P = \left(\frac{\partial H}{\partial T}\right)_P$$

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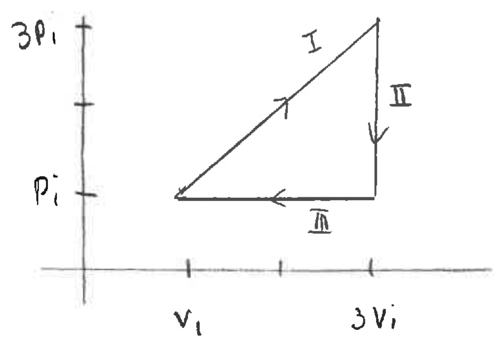
Question 3

A monoatomic ideal gas undergoes a cyclic process with the following stages, starting at initial pressure P_i and volume V_i. First it expands, during which $P = V$ at all times, until the volume reaches three times the initial volume. Second, the pressure drops back to the original pressure while the volume remains constant. Finally it is compressed back to its initial volume while the pressure remains constant.

a) Sketch the process on a PV diagram.



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b) Determine an expression for the work done in each stage.

1 $W = -\int P dV = -$ area under PV curve.

3 2 Stage I : $W = -\int_{V_i}^{3V_i} V dV$ since P = V

$= -\frac{V^2}{2}\Big|_{V_i}^{3V_i} = \frac{-9V_i^2 + V_i^2}{2} \Rightarrow W_I = -4V_i^2$

1 Stage II : $W = 0$ no area change $W_{II} = 0$

2 Stage III : $W = -\int_{3V_i}^{V_i} P_i dV = -P_i(-2V_i)$ $W_{III} = 2P_iV_i$

$= 2V_i^2$

Question 3 continued . . .

c) Determine an expression for the heat added in each stage.

1 $\left[Q = \Delta E - W \right.$

1 $\left[\Delta E = \Delta\left(\frac{3}{2} PV\right) = \frac{3}{2}\Delta(PV) \right.$

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Stage I $\quad \Delta(PV) = P_f V_f - P_i V_i$

$= 9V_i^2 - V_i^2 = 8V_i^2 \Rightarrow \Delta E = \frac{3}{2} 8V_i^2 = 12V_i^2$

Stage II $\quad \Delta(PV) = 3V_i \, \Delta P = 3V_i(-2V_i) = -6V_i^2 \Rightarrow \Delta E = -9V_i^2$

Stage III $\quad \Delta(PV) = P\Delta V = P_i(-2V_i) = -2V_i^2 \Rightarrow \Delta E = \frac{3}{2}(-2V_i^2)$

$= -3V_i^2$

Then:

	ΔE	W	Q
I	$12V_i^2$	$-4V_i^2$	$16V_i^2$
II	$-9V_i^2$	0	$-9V_i^2$
III	$-3V_i^2$	$2V_i^2$	$-5V_i^2$

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Question 4

An inventor claims to have produced a heat engine that uses a gas. In each cycle 1000 J of heat is supplied to the engine and the engine does 750 J of work. The lowest temperature during the engine cycle is 300 K.

a) Determine the efficiency of the engine.

$$\eta = \frac{W}{Q_{in}}$$

$$= 0.75$$

b) Determine the minimum value of the temperature of the high temperature reservoir needed to run this engine.

$$\eta \le 1 - \frac{T_c}{T_h}$$

$$\frac{T_c}{T_h} \le 1 - \eta$$

$$\frac{T_c}{1-\eta} \le T_h$$

$$\frac{300\,K}{0.25} \le T_h \Rightarrow T_h \geq 1200\,K$$

1200K is min

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Question 5

Answer either part a) or part b) for full credit for this problem.

a) The entropy of a monoatomic ideal gas is

$$S(T,V) = \frac{3}{2}Nk\ln(T) + Nk\ln(V) + g(N)$$

where $g(N)$ is an unknown function of N. This gas undergoes a free expansion. Determine whether ΔS is positive, negative or zero. Determine whether the change in Gibbs free energy is positive, negative or zero.

In a free expansion T is constant.

$$\Delta S = \frac{3}{2}Nk\ln(T_f) + Nk\ln(V_f) - \left[\frac{3}{2}Nk\ln(T_i) + Nk\ln V_i\right]$$

$$= \underbrace{\frac{3}{2}Nk\ln\left(\cancel{T_f/T_i}\right)}_{0} + Nk\ln\left(\frac{V_f}{V_i}\right)$$

$$\Rightarrow \Delta S = Nk\ln\left(\frac{V_f}{V_i}\right) \quad \text{but} \quad V_f > V_i \Rightarrow \Delta S > 0$$

$$G = E - TS + PV$$

$$dG = dE - TdS - SdT + PdV + VdP$$

$$= \cancel{TdS} - \cancel{PdV} - \cancel{TdS} - SdT + \cancel{PdV} + VdP$$

$$= -SdT + VdP$$

No change in T $\Rightarrow$ $\Delta G = \int V\,dP$

But V increases $\Rightarrow$ P drops $\Rightarrow$ $\Delta G < 0$

Question 5 continued ...

b) Use the Helmholtz free energy to show that, for a gas,

$$\left(\frac{\partial P}{\partial T}\right)_V = \left(\frac{\partial S}{\partial V}\right)_T.$$

$$F = E - TS$$

$$dF = dE - TdS - SdT$$

$$= \cancel{TdS} - PdV - \cancel{TdS} - SdT$$

$$= -SdT - PdV$$

So $\left(\frac{\partial F}{\partial T}\right)_V = -S$

$$\left(\frac{\partial F}{\partial V}\right)_T = -P$$

Then $\frac{\partial^2 F}{\partial V \partial T} = -\left(\frac{\partial S}{\partial V}\right)_T$ $\qquad$ $\frac{\partial^2 F}{\partial T \partial V} = -\left(\frac{\partial P}{\partial T}\right)_V$

So $\left(\frac{\partial S}{\partial V}\right)_T = \left(\frac{\partial P}{\partial T}\right)_V$

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Question 6

An ideal gas undergoes a compression at constant pressure.

a) Which of the following is true regarding the change in entropy of the gas?

i) $\Delta S_{\text{gas}} = 0$.
ii) $\Delta S_{\text{gas}} < 0$.
iii) $\Delta S_{\text{gas}} > 0$.

b) Which of the following is true regarding the magnitude of the change in entropy of the environment?

i) $|\Delta S_{\text{env}}| = |\Delta S_{\text{gas}}|$.
ii) $|\Delta S_{\text{env}}| < |\Delta S_{\text{gas}}|$.
iii) $|\Delta S_{\text{env}}| > |\Delta S_{\text{gas}}|$.

Briefly explain your answers.

During the compression $\Delta S_{gas} = C_p \int dT/T$

$$= C_p \underbrace{\ln\left(\frac{T_F}{T_i}\right)}_{\text{negative}}$$

$$\Rightarrow \Delta S_{gas} < 0$$

Then $\Delta S_{env} + \Delta S_{gas} \geq 0 \Rightarrow \Delta S_{env} \geq -\Delta S_{gas}$

$$= |\Delta S_{gas}|$$

Then at earlier stages T for the gas is higher than T for environment

So $dS_{env} > -dS_{gas}$

$$\Rightarrow |\Delta S_{env}| > |\Delta S_{gas}|$$

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